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Express your answer using a chemical symbol. Hg. Part A Arrange the following oxides in order of increasing acidity. Rank from least acidic to most acidic. To rank items as equivalent, overlap them. CaO, Al<sub>2</sub>O<sub>3</sub>, SiO<sub>2</sub>, CO<sub>2</sub>, P<sub>2</sub>O<sub>5</sub>, SO<sub>3</sub>. Part A barium oxide with water Express your answer as a chemical equation. Identify all of the phases in your answer.

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Image Transcriptionclose. 7:40 | 65% on.masteringchemistry.com :D <Chapter 15 - Attempt 1 Exercise 15.29 2 of 11 || Review | Constants | Periodic Table Part A Consider the following reactions and their respective equilibrium constants:  $\text{NO(g)} + \text{Br}_2\text{(g)} = \text{NOBr(g)}$   $K_p = 5.3$   $2\text{NO(g)} = \text{N}_2\text{(g)} + \text{O}_2\text{(g)}$   $K = 2.1 \times 10^{30}$  Use these reactions and their equilibrium constants to predict the equilibrium ...

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