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Chemistry Chapter 14 Notes. reaction rate. reaction rate equation. instantaneous rate. General equation. the rate at which reactants change into products over time. change in concentration/change in time. The reaction rate at a particular time as opposed to the avera.... aA + bB = cC +dD.

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Answers: (a)  $8.4 \times 10^{-7}$  M/s. (b)  $2.1 \times 10^{-7}$  M/s SAMPLE EXERCISE 14.3 continued The decomposition of N<sub>2</sub>O<sub>5</sub> proceeds according to the following equation: If the rate of decomposition of N<sub>2</sub>O<sub>5</sub> at a particular instant in a reaction vessel is  $4.2 \times 10^{-7}$  M/s, what is the rate of appearance of (a) NO<sub>2</sub>, (b) O<sub>2</sub>?

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**Semester 1 - AP Chemistry**

Equilibrium cannot be established between the liquid and the gas phase if the top is removed from the bottle because the system is not closed; one of the components of the equilibrium, the Br<sub>2</sub> vapor, would escape from the bottle until all liquid disappeared. Thus, more liquid would evaporate than can condense back from the gas phase to the liquid phase.

**Answer Key Chapter 13 - Chemistry 2e | OpenStax**

have a mass of 157amu, 14 have a mass of 158amu, 14 have a mass of 159amu, and the remainder have a mass of 160amu? Solve for % of 160amu isotope:  $100 - 2 - 12 - 14 = 58\%$   $(.02 \times 156) + (.12 \times 157) + (.14 \times 158) + (.14 \times 159) + (.58 \times 160) = 159.14$  amu 5.

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